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New Advances in the Kinetic Modeling of Thermal Oxidation of Epoxy-Diamine Networks

Xavier Colin[1]*, *Justine Delozanne*[2] *and Gurvan Moreau*[2]

[1]PIMM, Arts et Métiers Institute of Technology, CNRS, CNAM, HESAM University, Paris, France, [2]Safran Composites, A Platform of Safran Tech, Itteville, France

This article deals with the thermal oxidation mechanisms and kinetics of epoxy-diamine (EPO-DA) networks used as composite matrices reinforced with carbon fibers in the aeronautical field. The first part of this article is devoted to a detailed presentation of the new analytical kinetic model. The so-called "closed-loop" mechanistic scheme, developed in the last 3 decades in our laboratory in order to accurately describe the thermal oxidation kinetics of saturated hydrocarbon polymers, is recalled. Its main characteristics are also briefly recalled. Then, the system of differential equations derived from this oxidation mechanism is analytically solved without resorting to the usual simplifying assumptions that seriously degrade the reliability of all kinetic models. On the contrary, the generalization of the proportionalities observed between the steady concentrations of the different reactive species (i.e., hydroperoxides and alkyl and peroxy radicals) to the entire course of thermal oxidation gives a series of much sounder equations. From this basis, the kinetic model is completed by considering new structure/property relationships in order to predict the consequences of thermal oxidation on the thermomechanical properties, in particular on the glass transition temperature (T_g). To reach this second objective, the two main mechanisms responsible for the alteration of the macromolecular network structure are recalled: chain scissions and crosslinking. Like any other chemical species, their kinetics are directly expressed from the oxidation mechanistic scheme using the classical concepts of chemical kinetics. The second part of this article is devoted to the checking of the kinetic model reliability. It is shown that this latter accurately simulates the experimental curves of carbonyl build-up and T_g decrease versus time of exposure determined in our laboratory for three EPO-DA networks under study, exposed in a wide variety of thermal oxidative environments. The values determined by inverse solving method for the different model parameters are discussed and their temperature dependence are elucidated. Finally, an end-of-life criterion is proposed for predicting the lifetime of EPO-DA networks involving a predominant chain scission process.

Keywords: epoxy matrix, thermal oxidation, analytical kinetic model, chain scissions, glass transition temperature, lifetime prediction

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***Correspondence:**
Xavier Colin
xavier.colin@ensam.eu

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INTRODUCTION

Most of the composite material structures used in the civil aeronautical sector are composed of an epoxy or polyimide matrix reinforced with carbon fibers. Since the early 1980s, many studies have clearly shown that these materials can perish by matrix embrittlement induced by thermal oxidation when used in service in their glassy state (Alston, 1980; Kerr and Haskins, 1984; Street et al., 1988; Young and Chang, 1988; Scola and Vontell, 1991; Nam and Seferis, 1992; Salin and Seferis, 1993; Bowles et al., 1994; Skontorp et al., 1995; Parvatareddy et al., 1996; Tsotsis, 1998; Tsotsis and Lee, 1998; Colin et al., 1999; Tsotsis et al., 2001; Lafarie-Frenot and Rouquie, 2004; Colin et al., 2005a; Schoeppner et al., 2007; Tandon et al., 2009; Colin et al., 2011; Tandon and Pochiraju, 2011; Colin and Verdu, 2012a; Tandon, 2012). Thermal oxidation is limited to a superficial layer due to its kinetic control by oxygen diffusion (Gillen and Clough, 1989; Audouin et al., 1994). "Spontaneous" cracks can then be initiated in this superficial layer (even in the absence of external mechanical loading) due to both the development of a tensile stress gradient and the catastrophic fall of its fracture properties (Colin et al., 2005b; Colin et al., 2011; Colin and Verdu, 2012a; Colin and Verdu, 2012b), thus allowing oxygen penetration into deeper layers (Bowles et al., 1993; Meador et al., 1996; Colin et al., 2005b). Repeating this sequential scenario is expected to allow damage to propagate until the core of the composite material (Colin et al., 2005b) and failure to occur untimely.

Since the early 1980s, different approaches used for predicting the lifetime of composite materials gave a central role to kinetic modeling (Nelson, 1983; Bowles and Meyer, 1986; Bowles and Nowak, 1988; Salin and Seferis, 1996a; Salin and Seferis, 1996b; Cunningham and McManus, 1996; McManus and Cunningham, 1997; Crews and McManus, 1997; McManus et al., 2000; Colin et al., 2001a; Colin et al., 2001b; Colin et al., 2002; Decelle et al., 2003; Colin and Verdu, 2003; Colin et al., 2005a; Colin and Verdu, 2005; Tandon and Pochiraju, 2006; Tandon et al., 2006; Olivier et al., 2008; Olivier et al., 2009; Barjasteh et al., 2009; Barjasteh et al., 2011; Pochiraju et al., 2008; Pochiraju and Tandon, 2009; Lafarie-Frenot et al., 2010; Colin et al., 2011; Tandon and Pochiraju, 2011; Pochiraju, 2012; Cinquin et al., 2016; Colin et al., 2016; Colin et al., 2020; Colin et al., 2021). Within the scientific community, it was quickly accepted that all the empirical steps had to be eradicated in kinetic models, so that these latter could be efficient as well in simulation as in extrapolation. In particular, our laboratory played a major role in contributing this objective over the past 2 decades (Colin et al., 2001a; Colin et al., 2001b; Colin et al., 2002; Decelle et al., 2003; Colin and Verdu, 2003; Colin et al., 2005a; Colin and Verdu, 2005; Lafarie-Frenot et al., 2010; Cinquin et al., 2016; Colin et al., 2016; Colin et al., 2020; Colin et al., 2021).

To date, two versions of the kinetic model are fully operational for composites materials made of diamine cross-linked epoxy (EPO-DA) matrices. The "numerical version" allows solving the complete problem under consideration with, as only assumption, the uniqueness of the oxidation site (Colin et al., 2020). In EPO-DA matrices, this site corresponds to the C–H bond located in α position of a heteroatom (O or N) in oxy-methylene ($-O-CH_2-$), amino-methylene ($>N-CH_2-$) and methanol groups ($>CH-OH$). Indeed, this latter is characterized by a lower dissociation energy ($E_D \approx 376$ kJ mol^{-1}) compared to the aliphatic C–H bond in polymethylenic sequences ($E_D \approx 393$ kJ mol^{-1}) and the aromatic C–H bond ($E_D \approx 393$ kJ mol^{-1}) (Colin et al., 2011). It is commonly noted PH in both the oxidation mechanistic scheme and the corresponding kinetic model.

The "analytical version," in contrast, gives access to the oxidation kinetics only at low conversion ratios, i.e., when $[PH] \approx [PH]_0$, because it results from the generalization of observations in the steady state regime to the entire course of thermal oxidation (Colin et al., 2021). As embrittlement due to oxidation generally occurs for very low conversion ratios in polymers and, more particularly, in EPO-DA networks (Colin et al., 2011; Colin and Verdu, 2012a; Colin and Verdu, 2012b), the use of such an analytical kinetic model seems to largely be sufficient, not only for determining the corresponding critical oxidation events, but also for fully applying the classical methodology for lifetime prediction. It should be mentioned that the reliability this analytical kinetic model was recently demonstrated for two EPO-DA matrices considered for applications of composite structures at temperatures typically between 70 and 150°C in the civil aeronautical sector (Colin et al., 2021).

This complementary article aims at starting the generalization of this new analytical kinetic model to the whole family of EPO-DA matrices. After having recalled its theoretical foundations and structure, it will be shown that this kinetic model also successfully describes the thermal oxidation kinetics of a third EPO-DA network studied over the past decade in our laboratory (Terekhina et al., 2013; Cinquin et al., 2016; Colin et al., 2016; Colin et al., 2020). It should be pointed out that these three EPO-DA networks have very different glass transition temperatures, typically ranged between 158 and 263°C, thus allowing to study the impact of the molecular mobility on the thermal oxidation kinetics. The values determined by inverse solving method for the different model parameters will be discussed and their temperature dependence will be elucidated. In addition, the kinetic model will be completed by adding new structure/property relationships in order to predict the consequences of thermal oxidation on the thermomechanical properties. Finally, a peculiar attention will be paid to a possible end-of-life criterion in order to predict the lifetime of EPO-DA networks.

THEORY

Foundations of the Kinetic Model

The mechanistic scheme chosen for accurately describing the thermal oxidation of EPO-DA networks has been extensively detailed in previous publications, for instance in (Colin et al., 2020; Colin et al., 2021). It is composed of the following six reactions:

Initiation:

1b) $2POOH \rightarrow P^{\bullet} + PO_2^{\bullet}$ (k_{1b})

Propagation:

2) $P^{\bullet} + O_2 \rightarrow PO_2^{\bullet}$ (k_2)
3) $PO_2^{\bullet} + PH \rightarrow POOH + P^{\bullet}$ (k_3)

Terminations:

4) $P^{\bullet} + P^{\bullet} \rightarrow$ Inactive products (k_4)
5) $P^{\bullet} + PO_2^{\bullet} \rightarrow$ Inactive products (k_5)
6) $PO_2^{\bullet} + PO_2^{\bullet} \rightarrow$ Inactive products + O_2 (k_6)

Where PH, POOH, $P^{\bullet}$, and $PO_2^{\bullet}$ designate an oxidation site (i.e., C–H bond in oxy-methylene, amino-methylene or methanol group), an hydroperoxide, and alkyl and peroxy radicals, respectively. In addition, k_i (with i = 1b, . . . , 6) are rate constants.

The main peculiarity of this mechanistic scheme is to produce its own initiator. Indeed, radicals are formed by the thermal decomposition (according to the bimolecular mode) of the main propagation product: the hydroperoxide group (POOH). This closed-loop character explains well the sharp auto-acceleration of the oxidation reaction at the end of the induction period (Colin et al., 2006).

Recently, it was shown that the system of differential equations derived from the closed-loop mechanism can analytically be solved by generalizing the proportionality relationships observed between the steady concentrations of the different reactive species, i.e., $[POOH]$, $[PO_2^{\bullet}]$, and $[P^{\bullet}]$, to the entire course of thermal oxidation (Colin et al., 2021). These simplifying (but quite realistic) assumptions allowed obtaining an analytical kinetic model much sounder than all the other developed until now. In particular, this new kinetic model is able to accurately describe the first three stages of the thermal oxidation kinetics, i.e., the induction period, the auto-acceleration of the oxidation kinetics at the end of the induction period, and the steady-state regime. Unfortunately, due to the assumption of the low conversion ratios (i.e., $[PH] \approx [PH]_0 \approx$ constant), it cannot allow accounting for the sudden slow-down of the oxidation kinetics at long-term. This behavioral deviation between theory and experiment typically appears when $[PH]$ decreases by about ten percent. Then, it amplifies with exposure time.

The following solution was proposed for $[POOH]$ (Colin et al., 2021):

$$[POOH] = \frac{[POOH]_\infty}{1 + b\,Exp(-Kt)} \tag{1}$$

with:

$$K = k_3[PH]\left(\frac{k_{1b}}{k_6}\right)^{1/2}\left(\frac{\beta C}{1+\beta C}\right)^{1/2} \tag{2}$$

$$[POOH]_\infty = \frac{k_3[PH]}{2(k_{1b}\times k_6)^{1/2}}\left(\frac{\beta C}{1+\beta C}\right)^{1/2} \tag{3}$$

$$b = \frac{[POOH]_\infty - [POOH]_0}{[POOH]_0} \tag{4}$$

where $[POOH]_0$ and $[POOH]_\infty$ are the initial and steady concentrations of hydroperoxides, respectively.

Experimental measurements show that $[POOH]_\infty > 10\,[POOH]_0$ for weakly pre-oxidized polymer samples (Da Cruz et al., 2016; Huang et al., 2020). Consequently, $b >> 1$. Finally, the following order of magnitude ($b = 10$) was chosen for this study.

The following solutions were proposed for $[PO_2^{\bullet}]$ and $[P^{\bullet}]$ (Colin et al., 2021):

$$[PO_2^{\bullet}] = \frac{[PO_2^{\bullet}]_\infty}{1 + b\,Exp(-Kt)} \tag{5}$$

$$[P^{\bullet}] = [P^{\bullet}]_\infty\left[\frac{1}{1 + b\,Exp(-Kt)}\right]^2 \tag{6}$$

with:

$$[PO_2^{\bullet}]_\infty = \frac{k_3[PH]}{2k_6}\,\frac{\beta C}{1+\beta C} \tag{7}$$

$$[P^{\bullet}]_\infty = \frac{k_3[PH]}{2k_5}\,\frac{1}{1+\beta C} \tag{8}$$

where $[PO_2^{\bullet}]_\infty$ and $[P^{\bullet}]_\infty$ are the steady concentrations of peroxy and alkyl radicals, respectively.

In **Eqs 2**, **3**, **7**, **8**, C is the oxygen concentration in the EPO-DA network under consideration, which depends on the oxygen partial pressure P_{O2} in the exposure environment according to the common Henry's law:

$$C = S_{O2} \times P_{O2} \tag{9}$$

where S_{O2} is the coefficient of oxygen solubility in the EPO-DA network under consideration. The different values of S_{O2} reported in literature for EPO-DA networks were compiled in reference (Colin et al., 2020).

Two subfamilies of EPO-DA networks were clearly put in evidence:

1) Epoxies cross-linked with an aliphatic diamine hardener (e.g., Jeffamine D230, Jeffamine D400, Ancamine 2049, or IPDA) for which $S_{O2} \approx 5.1 \times 10^{-8}$ mol.L^{-1}.Pa^{-1} whatever the temperature,
2) Epoxies cross-linked with an aromatic diamine hardener (e.g., CAF, DDM, or DDS) for which $S_{O2} \approx 1.45 \times 10^{-7}$ mol.L^{-1}.Pa^{-1} whatever the temperature.

Only the second value will be used in this study because the three EPO-DA networks under study are cross-linked by CAF or DDS. In addition, β^{-1} roughly corresponds to the critical oxygen concentration C_C above which oxygen excess is reached:

$$\beta = \frac{3}{C_C} = \frac{k_6 k_2}{k_5 k_3[PH]} \tag{10}$$

Calculation of the Physico-Chemical Properties

Still using the classical concepts of the chemical kinetics, it is now possible to calculate, from the previous mathematical expressions of $[POOH]$, $[PO_2^{\bullet}]$, and $[P^{\bullet}]$, several key physico-chemical properties that can easily be checked experimentally. The mathematical

expressions of some of them were already reported in reference (Colin et al., 2021). New expressions will be added at the end of this section in order to then calculate the thermomechanical properties, in particular the glass transition temperature.

Oxygen Consumption

Undoubtedly, hydroperoxide concentration [*POOH*] and the oxygen consumption *Q* are the most relevant properties because their mathematical expressions can be deduced from the closed-loop mechanism without having to use additional adjustable parameter. Both properties can thus allow accurately determining the different rate constants k_i from the experimental data. The mathematical expression of *Q* was already reported in reference (Colin et al., 2021):

$$Q = \frac{r_\infty}{K}\left\{Kt + Ln\left[\frac{1 + b\,Exp(-Kt)}{1+b}\right] - \frac{1}{1 + b\,Exp(-Kt)} + \frac{1}{1+b}\right\} \tag{11}$$

with:

$$r_\infty = 2r_0\frac{\beta C}{1+\beta C}\left[1 - \frac{\beta C}{2(1+\beta C)}\right] \tag{12}$$

$$r_0 = \frac{k_3^2\,[PH]^2}{4k_6} \tag{13}$$

where r_∞ and r_0 are the steady oxidation rates in the general case and in oxygen excess, respectively.

Carbonyl Groups

Another key property, also often chosen to assess the rate constants k_i, is the carbonyl concentration [*P=O*]. In the closed-loop mechanism, these oxidation products are formed in the initiation (1b) and termination reactions (6) by specific chemical events involving radicals (e.g., β scission, disproportionation, etc.), which generally compete with many other chemical events, especially with hydrogen abstraction giving alcohols. Therefore, the calculation of [*P=O*] requires the use of additional adjustable parameters, namely formation yields. The mathematical expression of [*P=O*] was also reported in reference (Colin et al., 2021):

$$[P=O] = \frac{r_{CO\infty}}{K}\left\{Kt + Ln\left[\frac{1 + b\,Exp(-Kt)}{1+b}\right] - \frac{1}{1 + b\,Exp(-Kt)} + \frac{1}{1+b}\right\} \tag{14}$$

with:

$$r_{CO\infty} = r_0\frac{\beta C}{1+\beta C}\left(\gamma_{1CO} + \gamma_{6CO}\frac{\beta C}{1+\beta C}\right) \tag{15}$$

where $r_{CO\infty}$ is the steady rate of carbonyl build-up, and γ_{1CO} and γ_{6CO} are the formation yields of carbonyls in initiation (1b) and termination (6), respectively. In a recent publication (Colin et al., 2020), it was found that γ_{1CO} and γ_{6CO} are increasing functions of molecular mobility. The same orders of magnitude reported for these two yields in reference (Colin et al., 2020) were kept in this study (see **Table 2**).

Macromolecular Changes

The last key properties that will be detailed below are the concentrations of chain scissions (*S*) and crosslinking events (*X*) because these two macromolecular changes are known to be responsible for the changes in the thermomechanical properties (Colin et al., 2011; Colin and Verdu, 2012a; Colin and Verdu, 2012b). The most famous chain scission mechanism in literature is the β scission of alkoxyl radicals, which occurs during initiation (1b) in the closed-loop mechanistic scheme. In contrast, crosslinking usually takes place in terminations (4), (5), and (6) by the coupling of the different radical species. Their respective rates can be written as follows:

$$\frac{dS}{dt} = \gamma_{1S} \times k_{1b}[POOH]^2 \tag{16}$$

$$\frac{dX}{dt} = \gamma_{4X} \times k_4[P^\bullet]^2 + \gamma_{5X} \times k_5[P^\bullet][PO_2^\bullet] + \gamma_{6X} \times k_6[PO_2^\bullet]^2 \tag{17}$$

where γ_{1S}, γ_{4X}, γ_{5X}, and γ_{6X} are the yields in chain scissions in initiation (1b) and crosslinking in terminations (4), (5), and (6), respectively.

In a first approximation, to avoid having too many unknown parameters compared to our identification possibilities using the inverse solving method, it was assumed that crosslinking mainly occurs in oxygen excess, thus leading to peroxide bridges. Consequently, the crosslinking events in termination (4) and (5) were neglected: $\gamma_{4X} \approx \gamma_{5X} \approx 0$, and **Eq. 17** was simplified as:

$$\frac{dX}{dt} \approx \gamma_{6X} \times k_6[PO_2^\bullet]^2 \tag{18}$$

The replacement of [*POOH*] and [$PO_2^\bullet$] by their mathematical expressions in **Eqs 16**, **18** leads to:

$$\frac{dS}{dt} = \gamma_{1S} \times r_0\frac{\beta C}{1+\beta C}\left[\frac{1}{1 + b\,Exp(-Kt)}\right]^2 \tag{19}$$

$$\frac{dX}{dt} = \gamma_{6X} \times r_0\left(\frac{\beta C}{1+\beta C}\right)^2\left[\frac{1}{1 + b\,Exp(-Kt)}\right]^2 \tag{20}$$

The integration of **Eqs 19**, **20** with respect to time *t* gives mathematical expressions similar in form to *Q* and [*P=O*]:

$$S = \frac{r_{S\infty}}{K}\left\{Kt + Ln\left[\frac{1 + b\,Exp(-Kt)}{1+b}\right] - \frac{1}{1 + b\,Exp(-Kt)} + \frac{1}{1+b}\right\} \tag{21}$$

$$X = \frac{r_{X\infty}}{K}\left\{Kt + Ln\left[\frac{1 + b\,Exp(-Kt)}{1+b}\right] - \frac{1}{1 + b\,Exp(-Kt)} + \frac{1}{1+b}\right\} \tag{22}$$

with:

$$r_{S\infty} = \gamma_{1S} \times r_0\frac{\beta C}{1+\beta C} \tag{23}$$

$$r_{X\infty} = \gamma_{6X} \times r_0\left(\frac{\beta C}{1+\beta C}\right)^2 \tag{24}$$

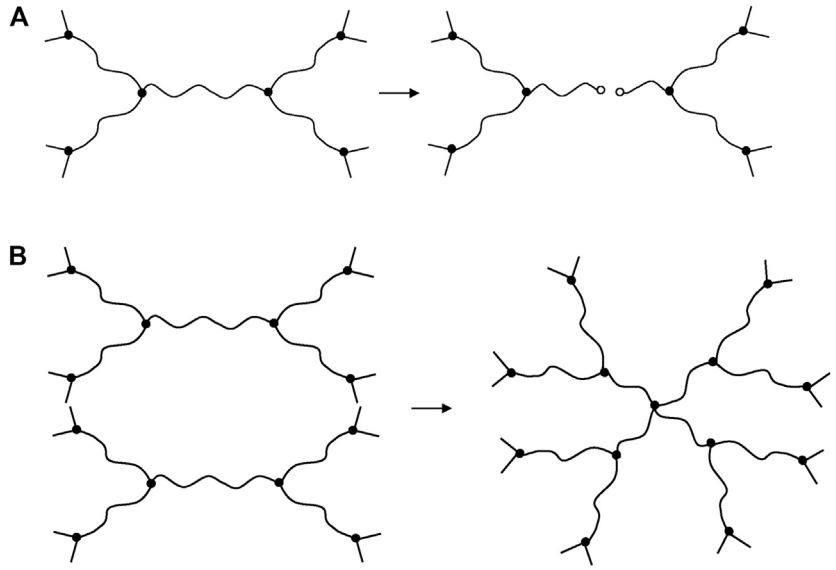

FIGURE 1 | Schematization of a chain scission **(A)** and a crosslinking event **(B)** in an ideal EPO-DA network.

where $r_{S\infty}$ and $r_{X\infty}$ are the steady rates of chain scissions and crosslinking, respectively.

Calculation of the Glass Transition Temperature

In an ideal EPO-DA network (i.e., without dangling chains), the ends of the elastically active chains are connected to tri-functional nodes. Consequently, each chain scission suppresses three elastically active chains and two nodes, and create two dangling chains, whereas each crosslinking event creates two new elastically active chains and one node of higher functionality (i.e., tetra-functional). All these macromolecular changes are schematized in **Figure 1**.

Thus, the concentrations of elastically active chains (ν) and nodes (n) can be written as a function of the concentrations of chain scissions (S) and crosslinking events (X) as follows:

$$\nu = \nu_0 - 3S + 2X \tag{25}$$

$$n = n_0 - 2S + X \tag{26}$$

where ν_0, n_0, ν and n are the concentrations of elastically active chains and nodes before and after thermal ageing, respectively.

The glass transition temperature T_g is an increasing function of n (or ν). It can thus be used to assess the impact of thermal oxidation on the macromolecular network. There is a lot of relationships between T_g and n (or ν) in literature, but to our opinion (Colin et al., 2011; Colin and Verdu, 2012a; Colin and Verbu, 2012b), the best one for polymer networks is the Di Marzio's equation (Di Marzio, 1964):

$$T_g = \frac{T_{gl}}{1 - K_{DM} \times F \times n} \tag{27}$$

where K_{DM} is an universal constant ($K_{DM} \approx 3$ for tri-functional crosslink nodes), F is the flex parameter characterizing the stiffness of the elastically active chains, and T_{gl} is the glass transition temperature of an hypothetical linear polymer containing all the structural units of the EPO-DA network under consideration, except its crosslink nodes.

In an ideal network (i.e., without dangling chains), the concentration of nodes is simply related to the concentration of elastically active chains by:

$$n = \frac{2}{f}\nu \tag{28}$$

where f is the node functionality (let us recall that $f = 3$ for ideal EPO-DA networks).

The introduction of **Eq. 28** into **Eq. 27** leads to:

$$T_g = \frac{T_{gl}}{1 - \frac{2}{3}K_{DM} \times F \times \nu} \tag{29}$$

Eq. 29 can be rewritten such as:

$$\frac{1}{T_g} = \frac{1}{T_{gl}} - \frac{2K_{DM} \times F \times \nu}{3T_{gl}} \tag{30}$$

which leads to:

$$\frac{1}{T_g} - \frac{1}{T_{g0}} = -\frac{2K_{DM} \times F}{3T_{gl}}(\nu - \nu_0) \tag{31}$$

where T_{g0} and T_g are the values of the glass transition temperature before and after thermal ageing, respectively.

The introduction of **Eq. 25** into **Eq. 31** leads to:

$$\frac{1}{T_g} - \frac{1}{T_{g0}} = \frac{2K_{DM} \times F}{3T_{gl}}(3S - 2X) \tag{32}$$

TABLE 1 | Molar mass of the repetitive constitutive unit (m_{CRU}), density (ρ), concentration of oxidation sites ([*PH*]) and glass transition temperature (T_g) for the three perfect EPO-DA networks under study.

Number	Network	m_{CRU} (g.mol^{-1})	ρ	(*PH*) (mol.L^{-1})	T_g (°C)
1	DGEBF-CAF	1,041	1.25	14.4	158 ± 2
2	DGEBA-CAF	1,097	1.25	13.7	182 ± 2
3	(*) Tactix 123/Tactix 742-DDS	1,674	1.11	10.6	263 ± 2

Comment:
**Matrix commercialized under the name Tactix by Hunstman.*
Abbreviations: CAF, 9,9-bis(3-chloro-4-aminophenyl)fluorene; DDS, 4,4'-Diamino diphenyl sulfone; DGEBA or Tactix 123, Diglycidyl ether of bisphenol A; DGEBF, Diglycidyl ether of bisphenol F; TGTPM or Tactix 742, Triglycidyl ether of triphenyl methane.

Thus, as expected, chain scissions decrease T_g whereas crosslinking increases it. The introduction of **Eqs 21**, **22** into **Eq. 32** leads finally to:

$$\frac{1}{T_g} = \frac{1}{T_{g0}} + \frac{2K_{DM} \times F}{3T_{gl} \times K}\left(3r_{S\infty} - 2r_{X\infty}\right)\left\{Kt + Ln\left[\frac{1 + b\,Exp\left(-Kt\right)}{1 + b}\right] - \frac{1}{1 + b\,Exp\left(-Kt\right)} + \frac{1}{1 + b}\right\} \tag{33}$$

EXPERIMENTAL

EPO-DA Networks

The three perfect EPO-DA networks (i.e., without dangling chains) under study are characterized by quite different T_g values, typically ranged between 158 and 263°C (see **Table 1**), thus allowing to analyze the possible effects of molecular mobility on the oxidation kinetics over a relatively wide temperature range. They result from the reaction of common bi-functional or tri-functional epoxy monomers with an aromatic amine hardener: 9,9-bis(3-chloro-4-aminophenyl) fluorine (CAF) or 4,4'-diamino diphenyl sulfone (DDS). Over the past decade, these three networks were considered in turn as potential matrices for composite structure applications at temperatures typically between 70 and 150°C in the civil aeronautical sector.

Network No. 3 was first considered in the early 2010s for potential applications at temperatures above 200°C, for which there is clearly a lack of solutions in organic matrix composites in Europe. Its thermal degradation was also the subject of a few articles, for instance (Terekhina et al., 2013; Cinquin et al., 2016; Colin et al., 2016). Networks No. 1 and 2 were only considered for applications near heat sources (i.e., aircraft engine applications) in the past 2 years. Their thermal degradation between 120 and 150°C is reported in recent articles (Colin et al., 2020; Colin et al., 2021).

The concentration in oxidation sites ([*PH*]) of these three perfect EPO-DA networks was directly determined from their theoretical repetitive monomer unit (see **Table 1**). It should then be pointed out that, when DDS is the amine hardener, the sulfonyl group protects the C–H bond of the amino-methylene group (i.e., $-CH_2-N<$) formed during the polymerization reaction against oxidation. Indeed, the sulfonyl group is a high electron-attracting group whose inductive effect through aromatic rings leads to an increase in the strength of the C–H bond. This stabilizing effect was first demonstrated by comparing the polymerization kinetics of different epoxy-diamine mixtures (Girard-Reydet et al., 1995). This is the reason why [*PH*] is lower for network 3 (about 10 mol.L^{-1}) compared to networks 1 and 2 (about 13–15 mol.L^{-1}).

Films and plates of these three EPO-DA networks with thicknesses typically ranging between 25 μm and 1 mm were produced by compression molding, then post-cured under primary vacuum (i.e., 10^{-3} bar) in accordance with the recommended industrial cure cycle in order to reach the maximum crosslinking density while avoiding any undesired pre-oxidation before thermal exposure. These precautions allowed minimizing the concentration of structural defects (in particular, in hydroperoxides POOH) in the EPO-DA networks. That is the reason why, as explained in *Foundations of the Kinetic Model*, a high value was chosen for the model parameter b in this study (typically, b = 10).

Thermal Ageing and Physico-Chemical Analyses

The oxidation kinetics of the three EPO-DA networks was studied at 120, 150, 180, and 200°C under an oxygen partial pressure ranged between 0.21 (i.e., ambient air) and 20 bars in autoclaves.

Films of EPO-DA networks were periodically removed from the autoclaves and cooled to room temperature in a desiccator containing silica gel for preventing any moisture recovery prior to be characterized by FTIR spectrophotometry. All FTIR spectra were recorded in a transmission mode between 400 and 4,000 cm^{-1} with a Perkin Elmer Frontier device, after having averaged the 16 scans obtained with a minimum resolution of 4 cm^{-1}.

As often reported in literature (Bellenger et al., 1981; Dyakonov et al., 1996; Rivaton et al., 1997; Colin et al., 2001a; Musto et al., 2001; Musto, 2003; Mailhot et al., 2005; Dao et al., 2006; Delor-Jestin et al., 2006; Longerias et al., 2007; Galant et al., 2010; Pei et al., 2011), the main structural changes were observed in the carbonyl region where two new wide absorption bands appeared and grew rapidly with exposure time: one centered around 1,690–1,670 cm^{-1}, and the other around 1,720–1,730 cm^{-1}. As an example, **Figure 2** shows the changes over time in the FTIR spectrum for network No. 3 at 150°C under 0.21 bar of oxygen (i.e., ambient air). These two bands were

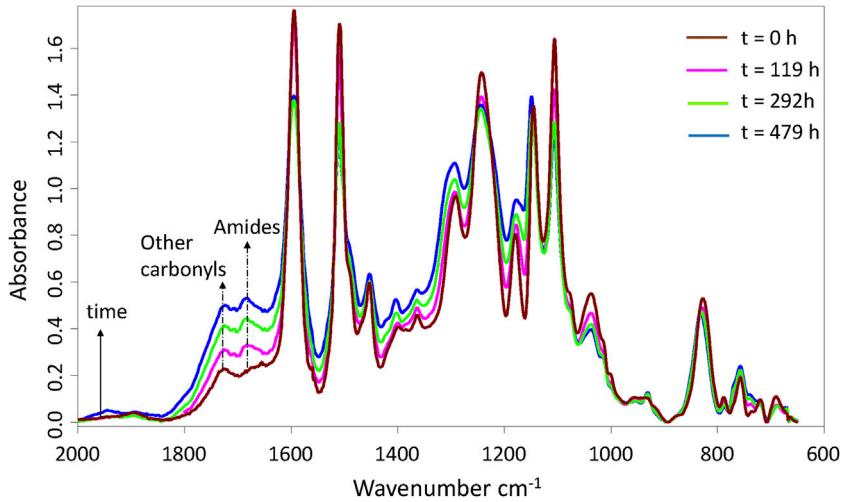

FIGURE 2 | Changes in the IR spectrum of network No. 3 (Tactix 123/Tactix 742-DDS) during its thermal ageing at 150°C under 0.21 bar of oxygen (i.e., ambient air).

assigned to amides and other types of carbonyl products, respectively.

Unfortunately, it was not possible to precisely identify these latter products mainly resulting from oxidation induced chain scissions in the hydroxyl propyl ether segment (Bellenger and Verdu, 1985) due to their wide variety: aldehydes, carboxylic acids, phenyl formates, etc ... Nevertheless, their average concentration throughout the film thickness [*P*=*O*] was determined by applying the common Beer-Lambert's law:

$$[P = O] = \frac{OD}{ep\,\varepsilon} \tag{34}$$

where *OD* is the optical density of the IR absorption band centered at 1720–1730 cm^{-1} (dimensionless), ε is the corresponding molar extinction coefficient (expressed in $L.mol^{-1}.cm^{-1}$), and *ep* is the film thickness (in cm).

Typical values of ε are ranged between 150 L $mol^{-1}.cm^{-1}$ (for ketones) and 850 L $mol^{-1}.cm^{-1}$ (for carboxylic acids) (Flett, 1962; Domke and Steinke, 1986; Lacoste and Carlsson, 1992; Barth, 2007; Da Cruz et al., 2016). In a first approximation, an average value of 500 L $mol^{-1}.cm^{-1}$ was chosen for carbonyl products in this study.

Plates of EPO-DA networks were removed from the autoclaves after definite durations and cooled at room temperature according to the same procedure as for films. Parallelepipedic shaped barrels of 25 × 2 × 1 mm^3 were machined from the plates in order to be analyzed by mechanical spectrometry in a tension mode between −110 and 330°C with a TA Instruments DMA Q800 device. The tests were performed with a controlled sinusoidal strain in the linear domain of the material viscoelasticity and the corresponding stress was measured. From the stress and strain values, the complex modulus E^* was calculated:

$$E^* = E' + j\,E'' \tag{35}$$

where E' is the storage (elastic component) modulus, E'' is the loss (viscous component) modulus and $tan(\delta) = E''/E'$ is the loss factor or damping, from which the phase shift δ between the stress and strain is extrapolated.

The frequency and heating rate were set at 1 Hz and 2°C min^{-1}, respectively. The principal relaxation temperature T_α, associated to the glass transition temperature T_g, was taken at the maximum of the α dissipation band. Examples of changes in the DMA thermogram of network No. 3 (Tactix 123/Tactix 742-DDS) are reported in **Supplementary Material**.

It should be mentioned that the repeatability of the aging tests was only checked for specific ageing conditions because of the long durations of these tests. As an example, **Figure 3** compares three different campaigns of ageing tests performed at 180°C under 0.21 bar of oxygen (i.e., ambient air). It can be seen that the three curves of carbonyl build-up only differ by the duration of their induction period, which is typically ranged between 6 and 13 h. Such a low experimental variability can be explained by small differences in the initial state of the samples (in particular, in their initial concentration of hydroperoxides) (Colin et al., 2006). In addition, at the end of the induction period, the three curves exhibit the same auto-acceleration of the oxidation kinetics.

Due to the high test repeatability, it was decided to group together all the experimental data obtained under the same ageing condition within a single experimental curve and no longer distinguish original from replicate data thereafter (both for [*P*=*O*] and T_g).

RESULTS AND DISCUSSION

Determination of Model Parameters

Let us recall that the validity of the analytical kinetic model constituted by **Eq. 1** to **Eq. 15** was already checked in a previous article devoted to the thermal oxidation kinetics of

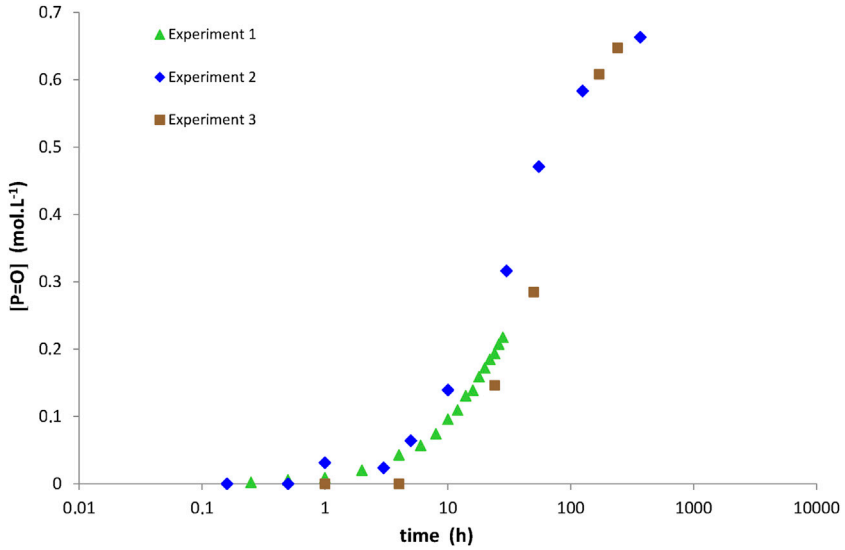

FIGURE 3 | Carbonyl build-up in 100 μm thick films of network No. 3 (Tactix 123/Tactix 742-DDS) at 180°C under 0.21 bar of oxygen (i.e., ambient air). Comparison between three different campaigns of ageing tests.

TABLE 2 | Values of parameters used for simulating the carbonyl build-up and T_g decrease for the three EPO-DA networks under study between 120 and 200°C with **Eq. 14** and **Eq. 33**, respectively.

Matrix	DGEBF-CAF (T_g = 158°C)		DGEBA-CAF (T_g = 182°C)		Tactix 123/Tactix 742-DDS (T_g = 263°C)			
T (°C)	120	150	120	150	120	150	180	200
k_{1b} (L.mol^{-1}.s^{-1})	2.0 10^{-3}	2.0 10^{-2}	2.0 10^{-3}	8.0 10^{-3}	2.0 10^{-3}	1.0 10^{-2}	8.0 10^{-2}	2.0 10^{-1}
r_0 (mol.L^{-1}.s^{-1})	4.5 10^{-5}	2.5 10^{-4}	6.0 10^{-5}	6.5 10^{-4}	2.2 10^{-5}	2.2 10^{-4}	1.6 10^{-3}	8.0 10^{-3}
β (L.mol^{-1})	13	40	15	18	6	3	2.5	1
k_{1b} (L.mol^{-1}.s^{-1})	2.0 10^{-3}	2.0 10^{-2}	2.0 10^{-3}	8.0 10^{-3}	2.0 10^{-3}	1.0 10^{-2}	8.0 10^{-2}	2.0 10^{-1}
γ_{1CO} (%)	30	30	30	30	10	10	20	30
γ_{6CO} (%)	30	30	30	30	10	10	20	30
γ_{1S} (%)	–	–	–	–	–	–	–	0.16
γ_{6X} (%)	–	–	–	–	–	–	–	0

networks No. 1 and 2 at 120 and 150°C under an oxygen partial pressure ranged between 0.21 (i.e., ambient air) and 10 bars (Colin et al., 2021). It was shown that **Eq. 14** allows accurately accounting for the three first stages of the thermal oxidation kinetics, i.e., the induction period, the sharp auto-acceleration of the oxidation reaction at the end of the induction period, and the steady-state regime. However, it cannot account for the last stage, i.e., the sudden slow-down of the oxidation reaction, when the concentration of oxidation sites is vanishing, because of the starting assumption of low conversion ratios (i.e., $[PH] \approx [PH]_0 \approx$ constant). It was also shown that this behavioural deviation between theory and experiment appears when $[PH]$ typically decreases by about ten percent, and it amplifies with exposure time. However, as the embrittlement of EPO-DA matrices is generally observed at low conversion ratios, i.e. during the induction period or the auto-acceleration of the oxidation reaction, it was concluded that the analytical kinetic model is largely sufficient for fully applying the classical methodology for lifetime prediction thereafter. The corresponding values of the model parameters determined for networks No. 1 and 2 by the inverse solving method are recalled in **Table 2**.

Examples of simulations with **Eq. 14** for network No. 3 at 150°C under an oxygen partial pressure ranged between 0.21 and 20 bars, but also between 120 and 200°C under an oxygen partial pressure of 0.21 bar (i.e., ambient air), are now reported in **Figures 4**, **5**, respectively. The same conclusions as for networks No. 1 and 2 can now be done for network No. 3. The corresponding values of the model parameters determined for network No. 3 by the inverse solving method are also listed in **Table 2**. Let us recall that the order of magnitudes of γ_{1CO} and γ_{6CO} come from a recent publication (Colin et al., 2020).

The values thus obtained for the three other model parameters (i.e., k_{1b}, r_0, and β) were plotted in Arrhenius graphs in order to elucidate their temperature dependence. The following conclusions can be drawn:

1) **Eq. 14** gives almost the same values of k_{1b} as those recently determined with the "numerical version" of the kinetic

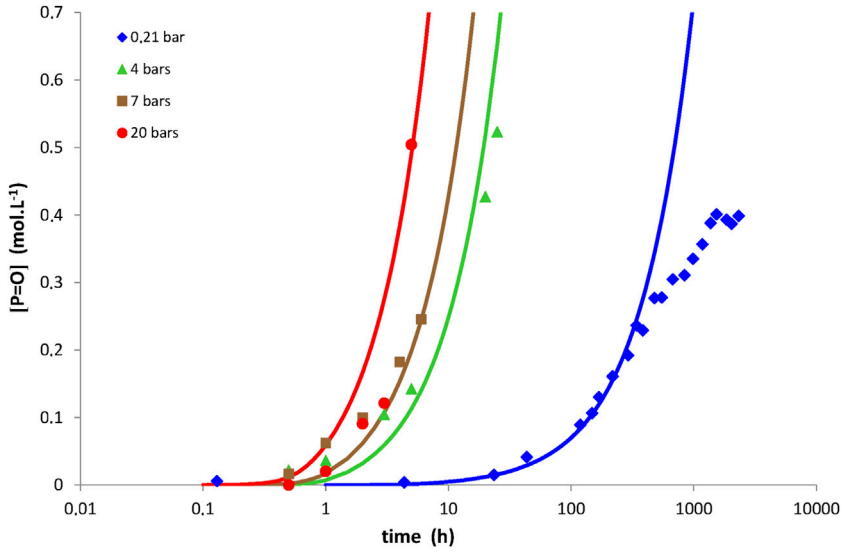

FIGURE 4 | Carbonyl build-up in 100 μm thick films of network No. 3 (Tactix 123/Tactix 742-DDS) at 150°C between 0.21 bar (i.e., ambient air) and 20 bars of oxygen. Comparison between simulations with **Eq. 14** (solid lines) and experimental data (points).

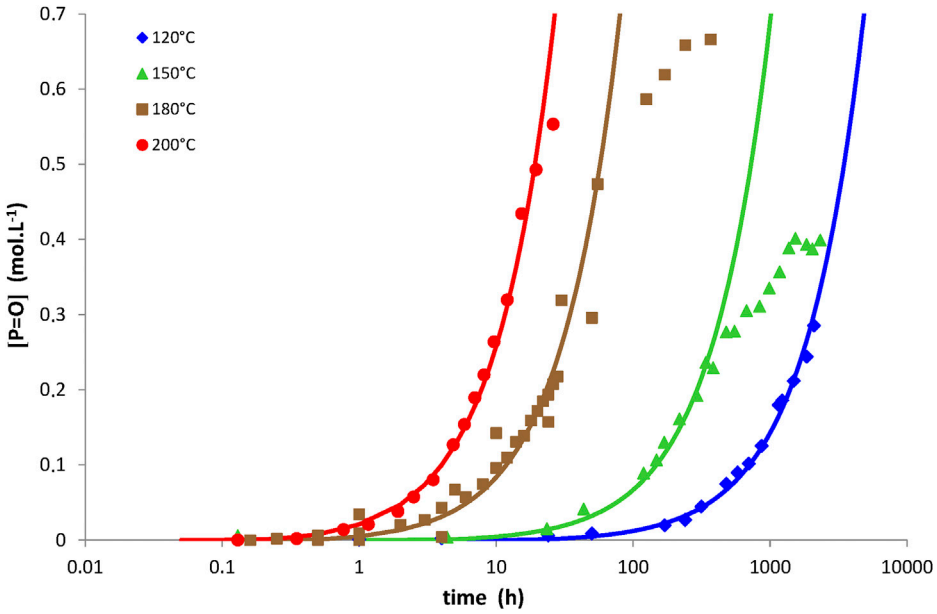

FIGURE 5 | Carbonyl build-up in 100 μm thick films of network No. 3 (Tactix 123/Tactix 742-DDS) between 120 and 200°C under 0.21 bar of oxygen (i.e., ambient air). Comparison between simulations with **Eq. 14** (solid lines) and experimental data (points).

model for all EPO-DA networks (Colin et al., 2020). Thus, it is confirmed that k_{1b} obeys the following general Arrhenius law:

$$k_{1b} = 1.9 \times 10^9 \times Exp - \frac{90\,000}{RT} L.mol^{-1}.s^{-1} \quad (36)$$

As an illustration, **Figure 6** reports the Arrhenius graph obtained for k_{1b} for networks No. 1, 2, and 3.

2) In contrast, the temperature dependence of parameters r_0 as β is much more complicated because they depend both on the concentration of oxidation sites (see **Eqs 10**, **13**, respectively) and the molecular mobility (Colin et al., 2020). It is therefore necessary to correct the values determined for these two parameters by the concentration of oxidation sites $[PH]$ and the effect of molecular mobility in order to obtain master curves in an Arrhenius diagram for all the EPO-DA

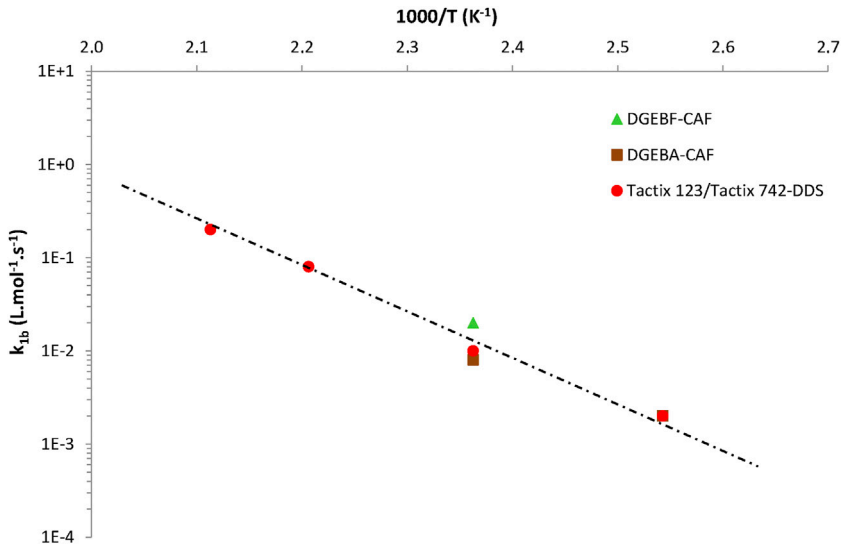

FIGURE 6 | Arrhenius graph between 120 and 200°C for the rate constant k_{1b} for the three EPO-DA networks under study.

networks. In a first approach, the impact of molecular mobility in EPO-DA networks was described taking a simple exponential function of T_g (Colin et al., 2020):

$$F(T_g) = Exp(T_g - 323) \tag{37}$$

Figures 7, **8** show the master curves obtained after correction for parameters r_0 and β, respectively. Unfortunately, because of the high T_g values of the three EPO-DA networks under study, almost only the glassy domain can be observed. However, a small jump of pre-exponential factor centered on the origin of the x-axis can be distinguished. In the glassy domain (i.e., for $T < T_g$), the following general Arrhenius laws can be proposed for parameters r_0 as β:

$$\frac{r_0}{[PH]^2 \times F^2(T_g)} = 4.7 \times 10^{-9} \times Exp - \left[\frac{133\,000}{R}\left(\frac{1}{T} - \frac{1}{T_g}\right)\right] \tag{38}$$

$$\frac{\beta}{[PH] \times F(T_g)} = 3.7 \times 10^{3} \times Exp - \left[\frac{39\,000}{R}\left(\frac{1}{T} - \frac{1}{T_g}\right)\right] \tag{39}$$

which leads to:

$$r_0 = 4.7 \times 10^{-9} \times [PH]^2 \times Exp\left[2\left(T_g - 323\right)\right] \times Exp - \left[\frac{133\,000}{R}\left(\frac{1}{T} - \frac{1}{T_g}\right)\right] \tag{40}$$

$$\beta = 3.7 \times 10^{3} \times [PH] \times Exp\left(T_g - 323\right) \times Exp - \left[\frac{39\,000}{R}\left(\frac{1}{T} - \frac{1}{T_g}\right)\right] \tag{41}$$

Let us recall that β^{-1} roughly corresponds to the critical oxygen concentration C_C above which oxygen excess is reached (see **Eq. 10**). Of course, the corresponding value of P_{O2C} can be easily deduced using the common Henry's law (**Eq. 9**). The values of C_C and P_{O2C} thus calculated for the three EPO-DA networks under study are reported in **Table 3**. In the glassy domain, it is found that C_C and P_{O2C} are both increasing functions of temperature. In the glass transition zone, typically at 150°C for network No. 1 (i.e., for DGEBF-CAF), it is found that P_{O2C} is about 5 bars, i.e., a value often reported for several other EPO-DA networks in literature, for instance in reference (Olivier 2008).

Prediction of Thermomechanical Properties

Then, **Eq. 33** was used to predict the changes in T_g of the different EPO-DA networks under study. Fortunately, the three parameters (i.e., K_{DM}, F, and T_{gl}) of the Di Marzio's equation were already carefully determined for network No. 3 in a previous study (Terekhina et al., 2013). Their values are recalled in **Table 4**.

Examples of simulations with **Eq. 33** for network No. 3 at 200°C under oxygen partial pressures of 0.21 bar (i.e., ambient air) and 10 bars are reported in **Figure 9**. It can be seen that **Eq. 33** satisfyingly predicts the catastrophic decrease in T_g because of a predominant chain scission process. In a first approach, the crosslinking events were totally ignored and thus, their yield γ_{6X} in the termination reactions was set to zero. It was found that a very small amount of chain scissions, calculated with a yield γ_{1S} much lower than unity (typically, γ_{1S} = 0.16%), allows explaining the observed changes in the thermomechanical behaviour of network No. 3.

It is common to define the temperature range of use of a composite material from the T_g value of its matrix. Indeed, it is generally considered that, when the temperature dangerously approaches T_g, the mechanical properties of the composite material (in particular, its elastic properties) start to drop dramatically. In the case of network No. 3, considered for high temperature applications in the aeronautic field because of its high initial T_g value (T_{g0} = 263°C), the fact that chain scissions largely predominate over crosslinking events poses a serious problem. Indeed, a too great decrease in T_g will no longer allow the composite material to fulfill its initial function of mechanical

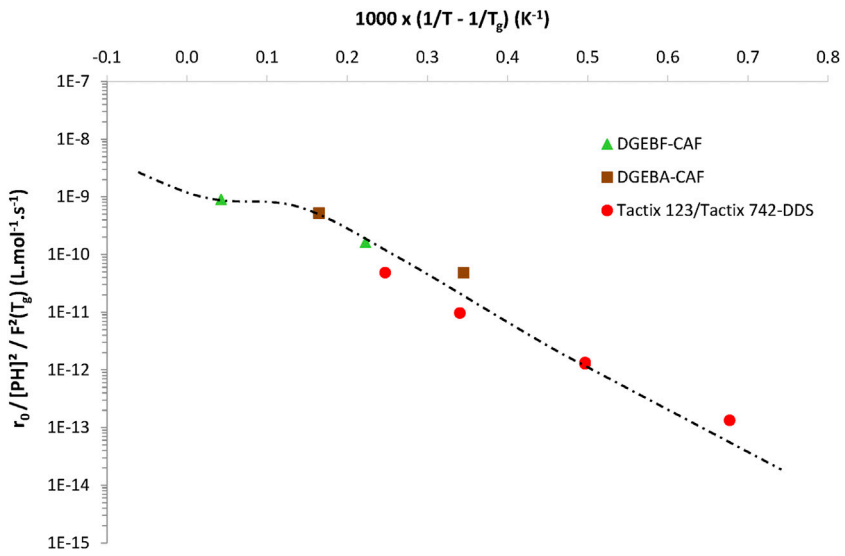

FIGURE 7 | Master curve for rate r_O for the three EPO-DA networks under study.

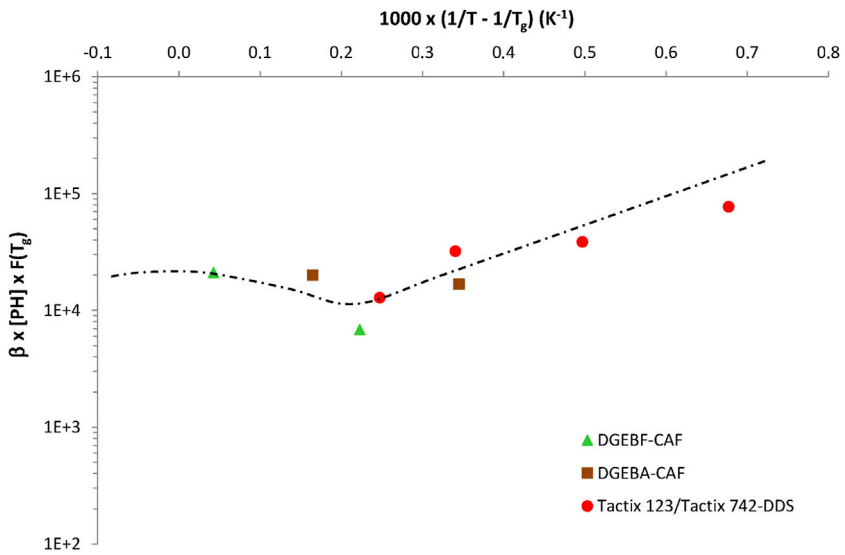

FIGURE 8 | Master curve for parameter β for the three EPO-DA networks under study.

TABLE 3 | Critical values of oxygen concentration and oxygen partial pressure for the three EPO-DA networks under study between 120 and 200°C.

Matrix	DGEBF-CAF (T_g = 158°C)		DGEBA-CAF (T_g = 182°C)		Tactix 123/Tactix 742-DDS (T_g = 263°C)			
T (°C)	120	150	120	150	120	150	180	200
C_C (mol.L^{-1})	2.3 10^{-1}	7.5 10^{-2}	2.0 10^{-1}	1.7 10^{-1}	5.0 10^{-1}	1.0	1.2	3.0
P_{O2C} (bar)	15.9	5.2	13.8	11.5	34.5	69.0	82.8	206.9

TABLE 4 | Values of the parameters of the Di Marzio's equation for network No. 3 (i.e., for Tactix 123/Tactix 742-DDS).

K_{DM}	F (g.mol.$^{-1}$)	T_{gl} (K)
3	31.8	353.7

structure. From this brief analysis, it can be concluded that an end-of-life criterion for network No. 3 would be a minimum value of T_g, denoted T_{gF}. In other words, the lifetime t_F of network No. 3 could be determined from **Eq. 33** as follows:

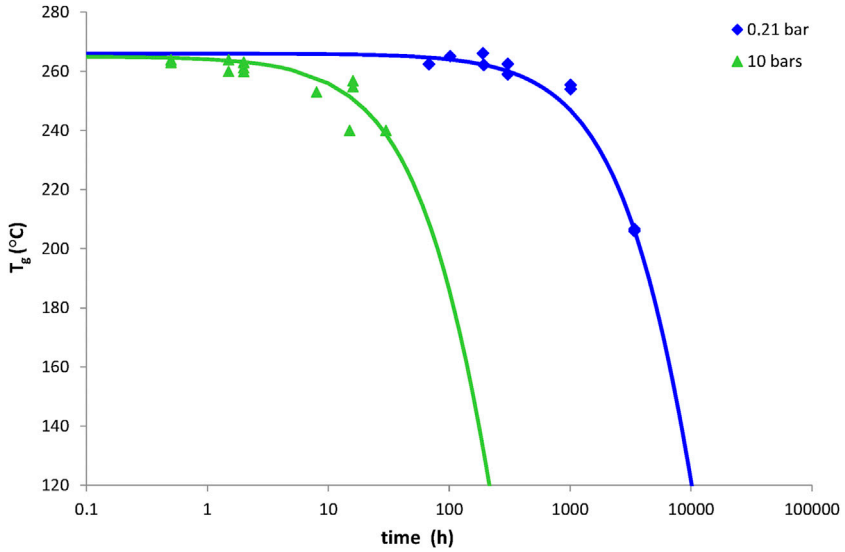

FIGURE 9 | T_g decrease in 1 mm thick plates of network No. 3 (Tactix 123/Tactix 742-DDS) at 150°C under 0.21 bar (i.e., ambient air) and 10 bars of oxygen. Comparison between simulations with **Eq. 33** (solid lines) and experimental data (points).

$$t = t_F \text{ when} T_g = T_{gF} \tag{42}$$

CONCLUSION

The analytical kinetic model developed for the thermal oxidation of two first EPO-DA networks in reference (Colin et al., 2021) was successfully generalized to an additional EPO-DA network in the present study. The model parameters were determined by inverse solving method from the experimental curves of carbonyl build-up and T_g decrease between 120 and 200°C under an oxygen partial pressure ranged between 0.21 (i.e., ambient air) and 10 bars. It was confirmed that the rate constant k_{1b} obeys a general Arrhenius law for all the EPO-DA networks. In contrast, the temperature dependence of parameters r_0 as β is much more complicated because they depend both on the concentration of oxidation sites and the molecular mobility. However, it was shown that master curves can be obtained in an Arrhenius diagram for these two parameters if their values are corrected by the concentration of oxidation sites and a relatively simple exponential function of T_g, proposed for describing the impact of molecular mobility.

In its current form, this new analytical kinetic model can be now easily implemented into commercial mechanical calculation codes for determining the consequences of thermal oxidation on the thermomechanical properties and predicting the lifetime of composite material structures. In particular, when chain scissions largely predominate over crosslinking, a minimum value of T_g can be used as end-of-life criterion.

DATA AVAILABILITY STATEMENT

The raw data supporting the conclusions of this article will be made available by the authors, without undue reservation.

AUTHOR CONTRIBUTIONS

XC–experimental testing and results post-processing, kinetic model development and calculation, writing original manuscript. JD–experiment scheme guidance, manuscript revision. GM–manuscript revision, project coordinator.

SUPPLEMENTARY MATERIAL

The Supplementary Material for this article can be found online at: https://www.frontiersin.org/articles/10.3389/fmats.2021.720455/full#supplementary-material

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